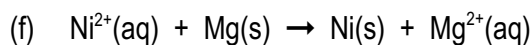
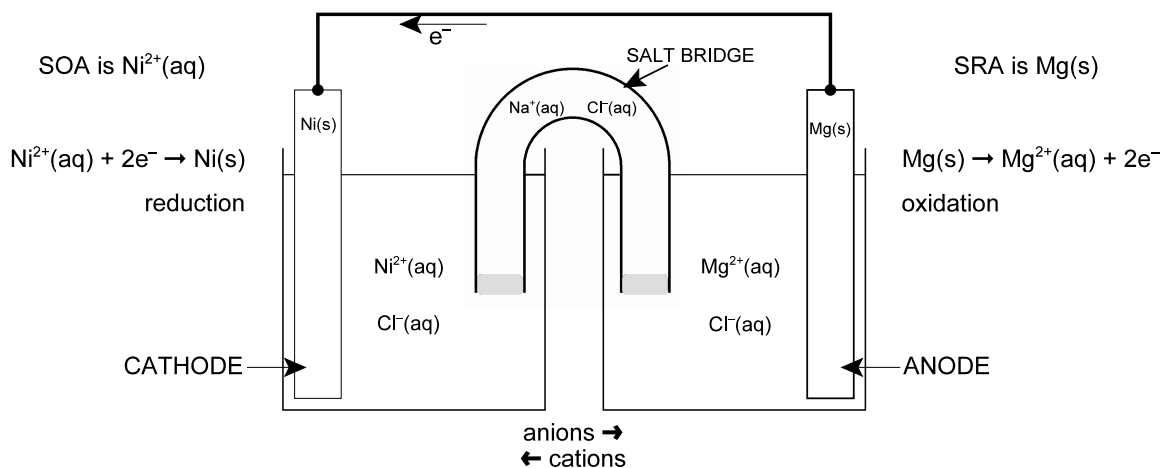


PRACTICE: GALVANIC CELLS  
ANSWERS

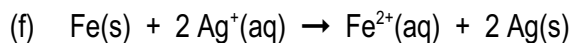
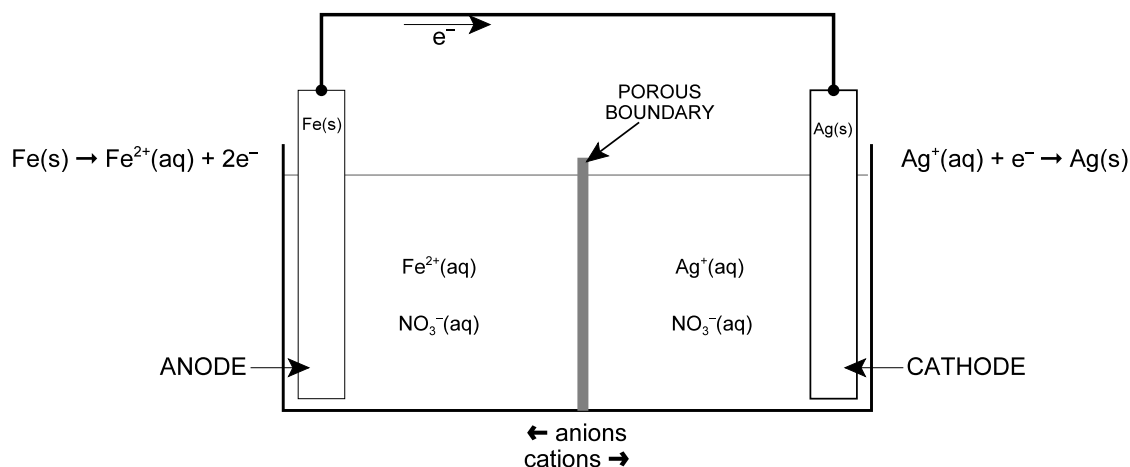
1. (a) to (e)



(g)  $\Delta E_{\text{cell}} = E_r(\text{cathode}) - E_r(\text{anode})$   
 $= (-0.26 \text{ V}) - (-2.37 \text{ V})$   
 $= +2.11 \text{ V}$



2. (a) to (e)



(g)  $\Delta E_{\text{cell}} = E_r(\text{cathode}) - E_r(\text{anode})$   
 $= (+0.80 \text{ V}) - (-0.44 \text{ V})$   
 $= +1.24 \text{ V}$